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Ethyl 3,3,3-trifluoro-2-hydroxy-2-(5-methoxy-1*H*-indol-3-yl)propionateZukhra Kadirova,^{a*} Samat Tolipov,^b Oleg Fedorovskiy,^c Ibragimov Bakhtiyar^b and Nusrat Parpiev^d

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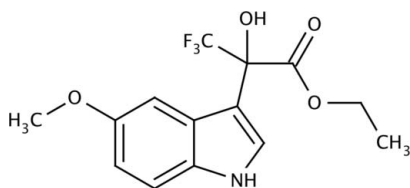
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}–\text{C}) = 0.003$ Å; R factor = 0.040; wR factor = 0.116; data-to-parameter ratio = 12.4.

In the title compound, $\text{C}_{14}\text{H}_{14}\text{F}_3\text{NO}_4$, the 3,3,3-trifluoropyruvate fragment has a *syn* configuration and is noncoplanar with the indole plane [dihedral angle = $84.87(5)^\circ$]. In the crystal, molecules form inversion-related dimers *via* pairs of intermolecular $\text{O}–\text{H}\cdots\text{O}$ hydrogen bonds. These dimers are connected by intermolecular $\text{N}–\text{H}\cdots\text{O}=\text{C}(\text{CF}_3)$ hydrogen bonds to form a two-dimensional network structure.

Related literature

For background on the synthesis and activity of trifluoropyruvates of indole, see: Nakamura *et al.* (2008); Abid *et al.* (2008). For the crystal structures of related compounds, see: Choudhury *et al.* (2004); Abid *et al.* (2008).



Experimental

Crystal data

 $\text{C}_{14}\text{H}_{14}\text{F}_3\text{NO}_4$ $M_r = 317.26$

Monoclinic, $P2_1/n$
 $a = 9.6277(4)$ Å
 $b = 15.9760(6)$ Å
 $c = 9.9738(4)$ Å
 $\beta = 109.314(5)^\circ$
 $V = 1447.75(10)$ Å³

$Z = 4$
 Cu $K\alpha$ radiation
 $\mu = 1.15$ mm⁻¹
 $T = 293$ K
 $0.55 \times 0.45 \times 0.40$ mm

Data collection

Oxford Diffraction Xcalibur Ruby
 CCD diffractometer
 Absorption correction: multi-scan
 (*CrysAlis PRO*; Oxford
 Diffraction, 2009)
 $T_{\min} = 0.782$, $T_{\max} = 1.000$

5738 measured reflections
 2911 independent reflections
 2319 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.018$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.040$
 $wR(F^2) = 0.116$
 $S = 1.06$
 2911 reflections
 234 parameters

H atoms treated by a mixture of
 independent and constrained
 refinement
 $\Delta\rho_{\max} = 0.20$ e Å⁻³
 $\Delta\rho_{\min} = -0.14$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D–H\cdots A$	$D–H$	$H\cdots A$	$D\cdots A$	$D–H\cdots A$
$\text{N1}–\text{H1A}\cdots\text{O2}^i$	0.86 (2)	2.11 (2)	2.9166 (18)	156.8 (18)
$\text{O2}–\text{H2B}\cdots\text{O3}^{ii}$	0.80 (2)	2.03 (2)	2.7798 (17)	156 (2)

Symmetry codes: (i) $x + \frac{1}{2}, -y + \frac{1}{2}, z + \frac{1}{2}$; (ii) $-x, -y, -z$.

Data collection: *CrysAlis PRO* (Oxford Diffraction, 2009); cell refinement: *CrysAlis PRO*; data reduction: *CrysAlis PRO*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *publCIF* (Westrip, 2010).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: PK2322).

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supplementary materials

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Ethyl 3,3,3-trifluoro-2-hydroxy-2-(5-methoxy-1*H*-indol-3-yl)propionate

Z. Kadirova, S. Tolipov, O. Fedorovskiy, I. Bakhtiyar and N. Parpiev

Comment

3,3,3-Trifluoropyruvates have been used as efficient fluorinated building blocks in the synthesis of some biologically active trifluoromethylated compounds owing to the unique properties of the trifluoromethyl group, such as high electronegativity, electron density, steric hindrance and its hydrophobic character. The incorporation of a tertiary α -trifluoromethyl alcohol stereocenter ($\text{CF}_3\text{C}^*(\text{OH})\text{R}_1\text{R}_2$) into heterocycles could provide novel drug candidates with unusual biological activities as a result of the presence of the chiral tertiary α -trifluoromethyl alcohol functionality (Nakamura *et al.* 2008, Abid *et al.*, 2008).

In this study we synthesized the 3,3,3-trifluoropyruvate derivative of indole, which is an analogue of indole alkaloids. The molecular structure is shown in Fig. 1.

The compound crystallizes as a racemate and the carboxy- and hydroxy- groups are *syn* to each other (torsion angle $\text{O2—C9—C10—O3} = 8.1$ (2)°). The 3,3,3-trifluoropyruvate fragment is non-coplanar to the plane of the indole [torsion angles C8—C7—C9—C10 , C6—C7—C9—C10 , C6—C7—C9—O2 , C8—C7—C9—O2 are 107.3 (2)°, -69.7 (2)°, 50.0 (2)°, and -133.0 (2)°, respectively]. The methoxy-, hydroxy- and trifluoromethyl groups deviate from the indole plane by 0.061 (2), 0.854 (1) and 0.165 (2) Å, respectively. The bond distances C—C and C—N in the indole group are in the range 1.369–1.435 Å and 1.358–1.379 Å, respectively. Due to the concurrent influence of electron-withdrawing groups in the 3,3,3-trifluoropyruvate fragment, the EtO(O)C—C(O)CF_3 bond is elongated slightly to 1.554 (2) Å.

The fluorine does not readily accept hydrogen bonds and hence behaves differently from chlorine and bromine, but a significant number of compounds pack *via* weak interactions involving organic fluorine (Choudhury *et al.*, 2004) and generate different packing motifs *via* $\text{F}\cdots\text{F}$, $\text{C—H}\cdots\text{F}$ and $\text{C—F}\cdots\pi$ interactions. In the presence of a strong acceptor such as C=O , the $\text{C—H}\cdots\text{O}$ interaction takes priority over $\text{C—H}\cdots\text{F}$. The molecules form inversion-related dimers *via* pairs of intermolecular $\text{O—H}\cdots\text{O}$ hydrogen bonds. These dimers are connected by intermolecular $\text{N—H}\cdots\text{O}=\text{C}(\text{CF}_3)$ hydrogen bonds to form a two-dimensional network structure.

Experimental

A solution of 0.170 g (0.001 mol) of ethyl 3,3,3-trifluoropyruvate in 10 ml of ether was added to a solution of 0.147 g (0.001 mol) of 5-methoxyindole in 10 ml ether, and stirred at room temperature for 24 h. The reaction was monitored by TLC. The reaction mixture was then evaporated under reduced pressure and the residue was purified by chromatography on silica gel (chloroform/ethylacetate=9:1); yield: 0.150 g (47%). The compound was crystallized from ethanolic solution by slow evaporation, giving colorless prism crystals suitable for X-ray diffraction analysis. ^1H NMR (300 MHz, dmsO-d_6): δ 1.11 (3*H*, t, CH_3); 4.22 (2*H*, q, CH_2 , $J_{\text{CH}_3\text{—CH}_2}$ 7,2 Hz); 3.72 (3*H*, s, O—CH_3), 6.78 (1*H*, d, aromatic, $J_{6-5} = 6.9$ Hz), 7.17 (1*H*, s, aromatic), 7.32 (1*H*, d, aromatic), 7.37 (2*H*, s. br., OH, CH), 11.58 (1*H*, s. br., NH); ^{19}F NMR (dmsO-d_6): 3.01 (s, CF_3); MS (m/z 317).

Refinement

Hydrogen atoms of the NH, OH and hydrogens attached to carbon (except C_{Me}) were located in difference Fourier maps and fully refined (including U_{iso}). Methyl hydrogens were included using a riding model with $U_{\text{iso}}(\text{H})$ values of $1.5 U_{\text{eq}}(\text{C}_{\text{Me}})$.

Figures

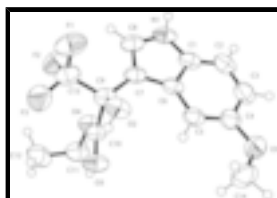


Fig. 1. A displacement ellipsoid plot drawn at the 50% probability level with H atoms shown as small spheres of arbitrary radius.

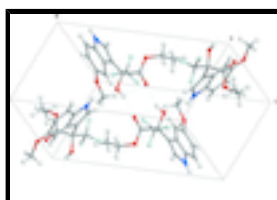


Fig. 2. A view of the crystal structure packing, showing part of the hydrogen bonding network.

Ethyl 3,3,3-trifluoro-2-hydroxy-2-(5-methoxy-1*H*-indol-3-yl)propionate

Crystal data

C₁₄H₁₄F₃NO₄

$M_r = 317.26$

Monoclinic, $P2_1/n$

Hall symbol: -P 2yn

$a = 9.6277$ (4) Å

$b = 15.9760$ (6) Å

$c = 9.9738$ (4) Å

$\beta = 109.314$ (5)°

$V = 1447.75$ (10) Å³

$Z = 4$

$F(000) = 656$

$D_x = 1.456$ Mg m⁻³

Cu $K\alpha$ radiation, $\lambda = 1.54178$ Å

Cell parameters from 2319 reflections

$\theta = 5.5\text{--}75.5^\circ$

$\mu = 1.15$ mm⁻¹

$T = 293$ K

Block, colourless

$0.55 \times 0.45 \times 0.40$ mm

Data collection

Oxford Diffraction Xcalibur Ruby CCD diffractometer

Radiation source: fine-focus sealed tube graphite

Detector resolution: 10.2576 pixels mm⁻¹

ω scans

Absorption correction: multi-scan (*Crys.Alis PRO*; Oxford Diffraction, 2009)

$T_{\text{min}} = 0.782$, $T_{\text{max}} = 1.000$

2911 independent reflections

2319 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.018$

$\theta_{\text{max}} = 75.9^\circ$, $\theta_{\text{min}} = 5.5^\circ$

$h = -11 \rightarrow 12$

$k = -19 \rightarrow 13$

$l = -12 \rightarrow 12$

5738 measured reflections

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.040$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.116$	$w = 1/[\sigma^2(F_o^2) + (0.0607P)^2 + 0.1918P]$
$S = 1.06$	where $P = (F_o^2 + 2F_c^2)/3$
2911 reflections	$(\Delta/\sigma)_{\max} < 0.001$
234 parameters	$\Delta\rho_{\max} = 0.20 \text{ e } \text{\AA}^{-3}$
0 restraints	$\Delta\rho_{\min} = -0.14 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: <i>SHELXL97</i> (Sheldrick, 2008), $F_c^* = kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
	Extinction coefficient: 0.0092 (10)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Bond distances, angles *etc.* have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell e.s.d.'s are taken into account in the estimation of distances, angles and torsion angles

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
F1	0.24919 (15)	0.21987 (8)	0.04081 (14)	0.0900 (4)
F2	0.24512 (15)	0.09413 (9)	-0.03770 (13)	0.0863 (4)
F3	0.42514 (12)	0.13538 (8)	0.14198 (14)	0.0805 (4)
O1	-0.16538 (14)	0.08409 (10)	0.51949 (16)	0.0773 (4)
O2	0.04297 (12)	0.12631 (7)	0.09810 (13)	0.0566 (3)
O3	0.12118 (13)	-0.03385 (7)	0.13702 (13)	0.0622 (3)
O4	0.34902 (11)	-0.00327 (7)	0.27903 (12)	0.0535 (3)
N1	0.34543 (16)	0.23072 (9)	0.50886 (17)	0.0590 (4)
C1	0.22216 (17)	0.19773 (10)	0.53026 (18)	0.0524 (4)
C2	0.1720 (2)	0.20300 (11)	0.6458 (2)	0.0627 (5)
C3	0.0439 (2)	0.16245 (13)	0.6367 (2)	0.0658 (5)

supplementary materials

C4	-0.03659 (18)	0.11812 (11)	0.5139 (2)	0.0596 (4)
C5	0.01333 (17)	0.11079 (10)	0.40028 (19)	0.0528 (4)
C6	0.14673 (16)	0.15055 (9)	0.40867 (17)	0.0476 (4)
C7	0.23112 (16)	0.15638 (9)	0.31430 (17)	0.0479 (4)
C8	0.34965 (18)	0.20619 (10)	0.3800 (2)	0.0556 (4)
C9	0.19239 (16)	0.11127 (9)	0.17494 (17)	0.0475 (4)
C10	0.21565 (15)	0.01523 (9)	0.19510 (16)	0.0461 (4)
C11	0.3900 (2)	-0.09195 (11)	0.2897 (2)	0.0639 (5)
C12	0.4403 (3)	-0.11618 (14)	0.1696 (3)	0.0826 (6)
H12A	0.4697	-0.1739	0.1794	0.124*
H12B	0.5223	-0.0818	0.1702	0.124*
H12C	0.3613	-0.1085	0.0817	0.124*
C13	0.2798 (2)	0.14066 (12)	0.0805 (2)	0.0645 (5)
C14	-0.2605 (2)	0.04673 (16)	0.3932 (2)	0.0810 (6)
H14A	-0.3506	0.0303	0.4073	0.122*
H14B	-0.2137	-0.0017	0.3703	0.122*
H14C	-0.2815	0.0863	0.3166	0.122*
H1A	0.403 (2)	0.2684 (13)	0.559 (2)	0.066 (6)*
H2A	0.222 (2)	0.2317 (15)	0.724 (2)	0.079 (6)*
H2B	0.009 (2)	0.0882 (16)	0.045 (2)	0.079 (7)*
H3A	0.007 (2)	0.1646 (14)	0.715 (2)	0.075 (6)*
H5A	-0.0408 (19)	0.0778 (12)	0.3176 (19)	0.056 (5)*
H8A	0.430 (2)	0.2217 (12)	0.3467 (19)	0.059 (5)*
H11A	0.304 (2)	-0.1242 (13)	0.292 (2)	0.068 (6)*
H11B	0.473 (3)	-0.0959 (14)	0.378 (2)	0.079 (6)*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
F1	0.1020 (9)	0.0653 (7)	0.0962 (9)	-0.0066 (6)	0.0240 (7)	0.0318 (6)
F2	0.0997 (9)	0.0956 (10)	0.0691 (7)	-0.0147 (7)	0.0352 (6)	-0.0033 (6)
F3	0.0560 (6)	0.0923 (9)	0.0963 (8)	-0.0088 (6)	0.0292 (6)	0.0143 (7)
O1	0.0549 (7)	0.0862 (10)	0.0964 (10)	-0.0112 (7)	0.0323 (7)	-0.0175 (8)
O2	0.0429 (6)	0.0470 (6)	0.0639 (7)	0.0081 (5)	-0.0037 (5)	-0.0038 (6)
O3	0.0518 (6)	0.0471 (6)	0.0710 (7)	-0.0071 (5)	-0.0022 (5)	0.0011 (5)
O4	0.0387 (5)	0.0429 (6)	0.0696 (7)	0.0059 (4)	0.0053 (5)	0.0037 (5)
N1	0.0492 (7)	0.0435 (7)	0.0687 (9)	-0.0106 (6)	-0.0014 (7)	-0.0051 (6)
C1	0.0448 (8)	0.0369 (7)	0.0645 (10)	0.0019 (6)	0.0032 (7)	-0.0040 (7)
C2	0.0611 (10)	0.0519 (10)	0.0645 (11)	0.0052 (8)	0.0066 (9)	-0.0157 (8)
C3	0.0611 (10)	0.0628 (11)	0.0724 (12)	0.0074 (9)	0.0204 (9)	-0.0139 (9)
C4	0.0469 (8)	0.0525 (9)	0.0774 (11)	0.0031 (7)	0.0179 (8)	-0.0085 (8)
C5	0.0420 (8)	0.0442 (8)	0.0648 (10)	-0.0021 (6)	0.0076 (7)	-0.0100 (7)
C6	0.0408 (7)	0.0331 (7)	0.0600 (9)	0.0026 (6)	0.0049 (6)	-0.0033 (6)
C7	0.0404 (7)	0.0358 (7)	0.0594 (9)	0.0005 (6)	0.0055 (6)	0.0032 (6)
C8	0.0437 (8)	0.0459 (9)	0.0673 (10)	-0.0053 (7)	0.0050 (7)	0.0068 (7)
C9	0.0365 (7)	0.0415 (8)	0.0563 (9)	0.0015 (6)	0.0042 (6)	0.0038 (6)
C10	0.0390 (7)	0.0426 (8)	0.0515 (8)	0.0001 (6)	0.0079 (6)	0.0010 (6)
C11	0.0521 (9)	0.0459 (9)	0.0860 (13)	0.0096 (8)	0.0125 (9)	0.0074 (9)

C12	0.0726 (13)	0.0660 (12)	0.1059 (17)	0.0106 (10)	0.0249 (12)	-0.0157 (12)
C13	0.0636 (10)	0.0580 (10)	0.0664 (11)	-0.0046 (8)	0.0139 (8)	0.0111 (9)
C14	0.0487 (10)	0.0883 (15)	0.1008 (16)	-0.0144 (10)	0.0176 (10)	-0.0043 (12)

Geometric parameters (Å, °)

F1—C13	1.330 (2)	C4—C5	1.375 (3)
F2—C13	1.339 (2)	C5—C6	1.410 (2)
F3—C13	1.332 (2)	C5—H5A	0.972 (18)
O1—C4	1.372 (2)	C6—C7	1.436 (2)
O1—C14	1.421 (3)	C7—C8	1.367 (2)
O2—C9	1.4090 (17)	C7—C9	1.499 (2)
O2—H2B	0.80 (2)	C8—H8A	0.974 (18)
O3—C10	1.1954 (18)	C9—C13	1.530 (3)
O4—C10	1.3137 (17)	C9—C10	1.554 (2)
O4—C11	1.465 (2)	C11—C12	1.484 (3)
N1—C8	1.357 (2)	C11—H11A	0.98 (2)
N1—C1	1.378 (2)	C11—H11B	0.98 (2)
N1—H1A	0.86 (2)	C12—H12A	0.9600
C1—C2	1.391 (3)	C12—H12B	0.9600
C1—C6	1.408 (2)	C12—H12C	0.9600
C2—C3	1.369 (3)	C14—H14A	0.9600
C2—H2A	0.90 (2)	C14—H14B	0.9600
C3—C4	1.405 (3)	C14—H14C	0.9600
C3—H3A	0.97 (2)		
C4—O1—C14	117.18 (16)	C7—C9—C13	113.86 (13)
C9—O2—H2B	110.4 (17)	O2—C9—C10	108.44 (12)
C10—O4—C11	116.54 (13)	C7—C9—C10	111.95 (13)
C8—N1—C1	109.48 (14)	C13—C9—C10	107.31 (14)
C8—N1—H1A	122.4 (13)	O3—C10—O4	126.01 (14)
C1—N1—H1A	127.1 (13)	O3—C10—C9	122.07 (13)
N1—C1—C2	131.13 (16)	O4—C10—C9	111.91 (12)
N1—C1—C6	107.25 (16)	O4—C11—C12	110.23 (17)
C2—C1—C6	121.60 (16)	O4—C11—H11A	107.4 (12)
C3—C2—C1	117.93 (17)	C12—C11—H11A	112.8 (12)
C3—C2—H2A	120.9 (14)	O4—C11—H11B	104.3 (13)
C1—C2—H2A	121.2 (14)	C12—C11—H11B	109.3 (13)
C2—C3—C4	121.43 (19)	H11A—C11—H11B	112.5 (17)
C2—C3—H3A	119.9 (13)	C11—C12—H12A	109.5
C4—C3—H3A	118.6 (13)	C11—C12—H12B	109.5
O1—C4—C5	124.53 (16)	H12A—C12—H12B	109.5
O1—C4—C3	114.21 (18)	C11—C12—H12C	109.5
C5—C4—C3	121.26 (17)	H12A—C12—H12C	109.5
C4—C5—C6	118.20 (15)	H12B—C12—H12C	109.5
C4—C5—H5A	120.7 (11)	F1—C13—F3	107.05 (16)
C6—C5—H5A	121.1 (11)	F1—C13—F2	107.46 (16)
C1—C6—C5	119.49 (16)	F3—C13—F2	106.77 (17)
C1—C6—C7	106.70 (14)	F1—C13—C9	111.24 (17)
C5—C6—C7	133.81 (15)	F3—C13—C9	113.91 (15)

supplementary materials

C8—C7—C6	106.71 (15)	F2—C13—C9	110.11 (15)
C8—C7—C9	129.49 (16)	O1—C14—H14A	109.5
C6—C7—C9	123.74 (13)	O1—C14—H14B	109.5
N1—C8—C7	109.86 (16)	H14A—C14—H14B	109.5
N1—C8—H8A	121.8 (11)	O1—C14—H14C	109.5
C7—C8—H8A	128.3 (11)	H14A—C14—H14C	109.5
O2—C9—C7	108.57 (13)	H14B—C14—H14C	109.5
O2—C9—C13	106.46 (13)		
C8—N1—C1—C2	178.10 (17)	C8—C7—C9—O2	-133.00 (17)
C8—N1—C1—C6	0.03 (18)	C6—C7—C9—O2	50.01 (18)
N1—C1—C2—C3	-179.49 (18)	C8—C7—C9—C13	-14.6 (2)
C6—C1—C2—C3	-1.7 (3)	C6—C7—C9—C13	168.40 (14)
C1—C2—C3—C4	-1.2 (3)	C8—C7—C9—C10	107.32 (18)
C14—O1—C4—C5	-7.2 (3)	C6—C7—C9—C10	-69.67 (18)
C14—O1—C4—C3	172.94 (18)	C11—O4—C10—O3	7.0 (3)
C2—C3—C4—O1	-177.34 (18)	C11—O4—C10—C9	-171.67 (15)
C2—C3—C4—C5	2.8 (3)	O2—C9—C10—O3	8.1 (2)
O1—C4—C5—C6	178.83 (16)	C7—C9—C10—O3	127.84 (17)
C3—C4—C5—C6	-1.3 (3)	C13—C9—C10—O3	-106.54 (18)
N1—C1—C6—C5	-178.63 (14)	O2—C9—C10—O4	-173.15 (13)
C2—C1—C6—C5	3.1 (2)	C7—C9—C10—O4	-53.40 (18)
N1—C1—C6—C7	0.41 (17)	C13—C9—C10—O4	72.22 (17)
C2—C1—C6—C7	-177.88 (14)	C10—O4—C11—C12	83.2 (2)
C4—C5—C6—C1	-1.5 (2)	O2—C9—C13—F1	54.98 (18)
C4—C5—C6—C7	179.74 (16)	C7—C9—C13—F1	-64.61 (19)
C1—C6—C7—C8	-0.69 (17)	C10—C9—C13—F1	170.92 (14)
C5—C6—C7—C8	178.15 (17)	O2—C9—C13—F3	176.06 (15)
C1—C6—C7—C9	176.89 (13)	C7—C9—C13—F3	56.5 (2)
C5—C6—C7—C9	-4.3 (3)	C10—C9—C13—F3	-67.99 (19)
C1—N1—C8—C7	-0.48 (19)	O2—C9—C13—F2	-64.04 (18)
C6—C7—C8—N1	0.73 (18)	C7—C9—C13—F2	176.37 (13)
C9—C7—C8—N1	-176.66 (14)	C10—C9—C13—F2	51.91 (18)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
N1—H1A \cdots O2 ⁱ	0.86 (2)	2.11 (2)	2.9166 (18)	156.8 (18)
O2—H2B \cdots O3 ⁱⁱ	0.80 (2)	2.03 (2)	2.7798 (17)	156 (2)

Symmetry codes: (i) $x+1/2, -y+1/2, z+1/2$; (ii) $-x, -y, -z$.

Fig. 1

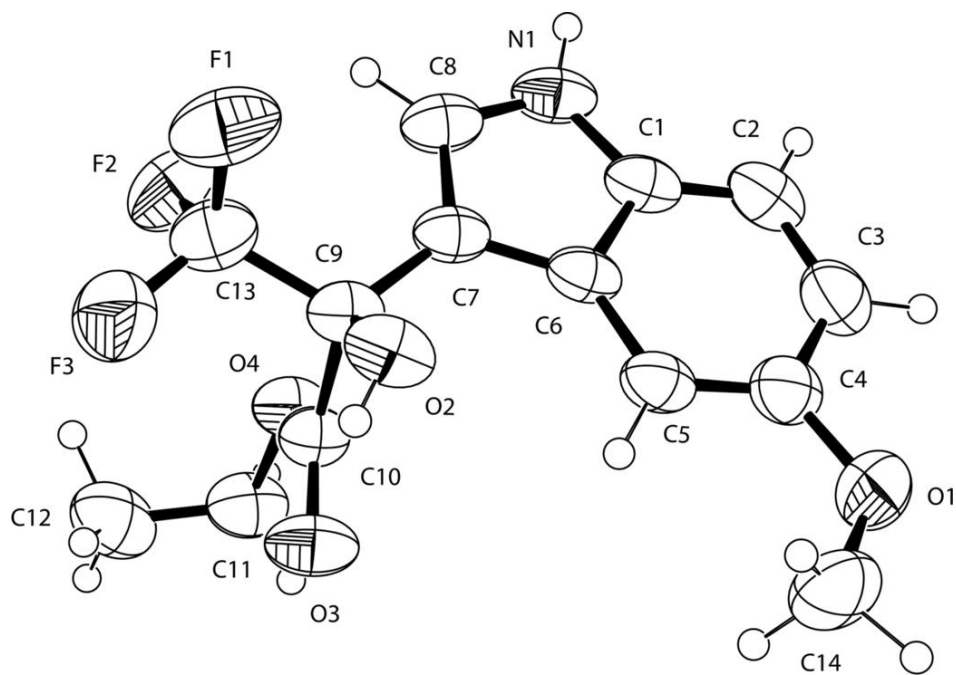


Fig. 2

